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on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

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Chapter 9: Stoichiometry Review and Chapter Summary ...

Stoichiometry. SECTION 1. SHORT ANSWER Answer the following questions in the space provided. 1. _____ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative

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number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

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SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual

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SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided

1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield

2 60 mol of N_2 are mixed with 120 mol of H

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Section 2 Worksheet

CHAPTER 9 REVIEW. Stoichiometry.
SECTION 9.2. PROBLEMS Write the
answer on the line to the left. Show all
your work in the space provided. 1. The
following equation represents a
laboratory preparation for oxygen gas:
 $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many
grams of O_2 form if 3.0 mol of KClO_3 are

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totally consumed? 2. Given the following equation ...

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the left show all your work in the space
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reaction is 22 g

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Show all your work in the space
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represents a laboratory preparation for
oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$... CHAPTER 9

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Reaction stoichiometry uses molar relationships to determine the amounts of unknown reactants or products from the amounts of known reactants or products. CHAPTER 9 DO NOT

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1.43 g/L calculate the volume of this gas. 76
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— . 81 g 6. A car air bag requires 70. L of
nitrogen gas ...

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Prentice Hall, Inc. All rights In your notebook, solve the following problems.

SECTION 9.1 THE ARITHMETIC OF

EQUATIONS Use the 3-step problem-

solving approach you learned in Chapter

4. 1. An apple pie needs 10 large apples,

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2 crusts (top and bottom), and 1
tablespoon of cinnamon.

9 Stoichiometry Practice Problems

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theoretical yield is 25 g Calculate the percentage yield 2 60 mol of N_2 are mixed with 120 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g)$

Download Chapter 9 Review Stoichiometry Section 2 Work

Chapter 9 focuses on reaction stoichiometry: using a balanced

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chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

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CHAPTER 9 REVIEW Stoichiometry
SECTION 2 PROBLEMS Write the answer
on the line to the left. Show all your work
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 $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many

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moles of O_2 form if 3.0 mol of $KClO_3$
are totally consumed? ...

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