

## Chemistry Review Solutions For 2008

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AP\* Chemistry: 2008 Released Multiple Choice Exam NO CALCULATORS MAY BE USED Note: For all questions, assume that the temperature is 298 K, the pressure is 1.00 atmosphere, and solutions are aqueous unless otherwise specified. Throughout the test the following symbols have the definitions specified unless otherwise noted. Part A

### AP\* Chemistry: 2008 Released Multiple Choice Exam

A 50.0 mL solution of 0.150 M HCl is mixed with 25.0 mL of 0.400 M HCl. What is the HCl concentration in the final solution? (Assume volumes are additive.) (A) 0.0175 (B) 0.233 (C) 0.275 (D) 0.550 11. For the reaction:  $2X + 3Y \rightarrow 3Z$ , the combination of 2.00 moles of X with 2.00 moles of Y produces 1.75 moles of Z.

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Assuming the density of water to be 1.0 g/mL the above solution contains 1.0 mg of Cl-in 1000 mL (1 L) of water so 1 ppm = 1 mg/L (or 1  $\mu\text{g/mL}$ ). The above assumption about the density of water becomes less valid at high temperatures or in highly concentrated solutions such as effluents or in seawater where

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