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Name Date Pd Chemistry - Unit 1  
Review 1. What is the difference between mass and volume? Mass is the amount of matter or “stuff” in a sample; volume is the space that sample occupies. 2. If the box at left contains atoms of aluminum in the liquid phase, represent the same atoms in the solid phase in the box at right. 3.

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Review Quiz - Quizizz a.  $100 \text{ cg} = 1 \text{ g}$  c.  
 $1 \text{ cm} = 1 \text{ mL}$  b.  $1000 \text{ mm} = 1 \text{ m}$  d.  $10 \text{ kg}$   
 $= 1 \text{ g}$  18. A cubic meter is about the  
same as the volume occupied by a \_\_\_\_.  
a. kilogram of water c. washing machine  
b. cup of milk d. basketball arena 19.  
The metric unit for

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One possible answer is that with both areas of study, students slowly build up a vocabulary and a set of concepts that relate the new terms. Both areas rely on a systematic approach—grammar versus the scientific method. Chemistry deals with matter and words do not qualify as matter.

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a.  $100 \text{ cg} = 1 \text{ g}$  c.  $1 \text{ cm} = 1 \text{ mL}$  b.  $1000 \text{ mm} = 1 \text{ m}$  d.  $10 \text{ kg} = 1 \text{ g}$  18. A cubic meter is about the same as the volume occupied by a \_\_\_\_.

a. kilogram of water  
c. washing machine  
b. cup of milk  
d. basketball arena

19. The metric unit for length that is closest to the thickness of a dime is the.

a. micrometer  
c. centimeter  
b. millimeter  
d. decimeter

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7th grade. 597 times. Chemistry. 73%  
average accuracy. a year ago.  
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the sugar and lemonade what method  
should you use? answer ...

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= 1 g c. 1 cm = 1 mL b. 1000 mm = 1 m  
d. 10 ...

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1.35 micrometers =  $1.35 \times 10^3$   
millimeters ? 0.079 mL = 0.079 cm<sup>3</sup> ?  
 $7.43 \times 10^{-2}$  nanoliters =  $7.43 \times 10^{-8}$   
milliliters ? 3.5 kilograms =  $3.5 \times 10^6$   
milligrams ?  $1.89 \times 10^4$  deciliters =  $1.89$   
 $\times 10^3$  liters

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Substances; 1.4...

## **AP Chemistry: 1.1-1.4 Moles, Mass Spectrometry, Elemental ...**

Chemistry - Unit 1 - Worksheet 6.

Dimensional Analysis. Use the factor-label method to make the following conversions. Remember to use the appropriate number of sf's in your answer. Part 1 . 1. 74 cm x = meters. 2.  $8.32 \times 10^{-2}$  kg x = grams. 3. 55.5 mL x = cm<sup>3</sup>. 4. 0.00527 cal x = kilocalories

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Review and study resources for the properties of macromolecules. Works great as an AP Biology Test Review or Study Guide as students work though Unit 1. This resource is designed to help guide your AP Biology students to learn, study, and review all of the topics in Unit 1. Perfect for distance learning or in class!

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Honors Chemistry Unit 1 Notes: Matter and Measurements. Objectives: 1. Students will become familiar with the meanings of various introductory concepts such as the meanings of the words: chemistry, substance, element, compound, atom, molecule, heterogeneous and homogeneous mixtures, and matter. ... (some are review) 1. Chemistry—definition ...

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