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pH of 10.11. What is the K_b for this base? $K_b = 1.1 \times 10^{-7}$ Q12. Draw the Lewis structures for HNO_3 and HNO_2 and explain which acid is stronger and why. .

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1. Bronsted-Lowry

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definition of acids and bases, pH scale, strong acids and bases.

Arrhenius' definition of acids and bases: Acid:

Substance that produces aqueous H^+ (protons) when dissolved in water.

Hydrochloric acid: $HCl(aq) \rightarrow H^+(aq) + Cl^-(aq)$ K_{eq} is large; strong acid.

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1. Bronsted-Lowry definition of acids and bases, pH scale, strong acids and bases.

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