

## Net Ionic Equation Worksheet With Answers

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### Net Ionic Equation Worksheet With

Net Ionic Equation Worksheet Answers Write balanced molecular, ionic, and net ionic equations (NIE) for each of the following reactions. Assume all reactions occur in aqueous solution. 1.  $2\text{NaCl}(\text{aq}) + \text{Pb}(\text{NO}_3)_2(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{NaNO}_3(\text{aq})$  Ionic Equation:  $2\text{Na}^+(\text{aq}) + 2\text{Cl}^-(\text{aq}) + \text{Pb}^{2+}(\text{aq}) + 2\text{NO}_3^-(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{Na}^+(\text{aq}) + 2\text{NO}_3^-(\text{aq})$

### Net Ionic Equation Worksheet Answers

equation is known as the net ionic equation. For example: The reaction of potassium chloride and lead II nitrate Molecular Equation:  $2\text{KCl}(\text{aq}) + \text{Pb}(\text{NO}_3)_2(\text{aq}) \rightarrow 2\text{KNO}_3(\text{aq}) + \text{PbCl}_2(\text{s})$  Complete Ionic Equation:  $2\text{K}^+(\text{aq}) + 2\text{Cl}^-(\text{aq}) + \text{Pb}^{2+}(\text{aq}) + 2\text{NO}_3^-(\text{aq}) \rightarrow 2\text{K}^+(\text{aq}) + 2\text{NO}_3^-(\text{aq}) + \text{PbCl}_2(\text{s})$  Net Ionic Equation:  $2\text{Cl}^-(\text{aq}) + \text{Pb}^{2+}(\text{aq}) \rightarrow \text{PbCl}_2(\text{s})$

### Net Ionic Equation Worksheet Answers

Net Ionic Equation Worksheet - answers 1. Molecular Equation:  $2\text{NaCl}(\text{aq}) + \text{Pb}(\text{NO}_3)_2(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{NaNO}_3(\text{aq})$  Ionic Equation:  $2\text{Na}^+(\text{aq}) + 2\text{Cl}^-(\text{aq}) + \text{Pb}^{2+}(\text{aq}) + 2\text{NO}_3^-(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{Na}^+(\text{aq}) + 2\text{NO}_3^-(\text{aq})$  NIE:  $2\text{Cl}^-(\text{aq}) + \text{Pb}^{2+}(\text{aq}) \rightarrow \text{PbCl}_2(\text{s})$  2. Molecular Equation:  $\text{Na}_2\text{CO}_3(\text{aq}) + \text{FeCl}_2(\text{aq}) \rightarrow \text{FeCO}_3(\text{s}) + 2\text{NaCl}(\text{aq})$  Ionic Equation:  $2\text{Na}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) + \text{Fe}^{2+}(\text{aq}) + 2\text{Cl}^-(\text{aq}) \rightarrow \text{FeCO}_3(\text{s}) + 2\text{Na}^+(\text{aq}) + 2\text{Cl}^-(\text{aq})$  NIE:  $\text{Fe}^{2+}(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{FeCO}_3(\text{s})$

### Net Ionic Equation Worksheet - answers

PRACTICE PROBLEMS ON NET IONIC EQUATIONS page 1 of 3 Show the complete ionic and net ionic forms of the following equations. If all species are spectator ions, please indicate that no reaction takes place. Note: you need to make sure the original equation is balanced before proceeding! A set of solubility rules are given at the end of this ...

### PRACTICE PROBLEMS ON NET IONIC EQUATIONS

Write the molecular, total ionic and net ionic equations for the following word equations, where possible: 1. acetic acid + calcium hydroxide → calcium acetate + water a) molecular eq'n b) total ionic eq'n c) net ionic eq'n 2. sulphuric acid + barium nitrate → barium sulphate + nitric acid a) molecular eq'n b) total ionic eq'n

### Net ionic equations worksheet - Ms. Mogck's Classroom

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### Net Ionic Equation Worksheets - Learn Kids

Steps for Writing Net Ionic Equations Determine if each compound is soluble (dissolved) or insoluble (solid) in water using the solubility rules.  $\text{AgNO}_3(\text{aq}) + \text{NaCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{NaNO}_3(\text{aq})$  (Rule number 4 says that chlorides are soluble except when paired with Silver, Lead, or Mercury)

### Net Ionic Equation Worksheet - TeacherPage

Ionic Equation Worksheet Write balanced molecular, total ionic, and net ionic equations for each of the following. 1. Aqueous sodium hydroxide reacts with aqueous copper(II) sulfate to precipitate copper(II) hydroxide. 2. Aqueous potassium carbonate reacts with aqueous silver nitrate to precipitate silver carbonate. 3.

### Ionic Equation Worksheet - Mohamad Berry

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Net Ionic Equation Worksheet Answers. Write balanced molecular, ionic, and net ionic equations (NIE) for each of the following reactions. Assume all reactions occur in aqueous solution. 1.  $2\text{NaCl}(\text{aq}) + \text{Pb}(\text{NO}_3)_2(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{NaNO}_3(\text{aq})$  Ionic Equation:  $2\text{Na}^+(\text{aq}) + 2\text{Cl}^-(\text{aq}) + \text{Pb}^{2+}(\text{aq}) + 2\text{NO}_3^-(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{Na}^+(\text{aq}) + 2\text{NO}_3^-(\text{aq})$

### Net Ionic Equation Worksheet Answers

Net ionic equation worksheet answers. Aqueous solutions of ammonium phosphate and sodium sulfate are mixed. When two solutions of ionic compounds are mixed a solid may form. Net ionic equation worksheet answers write balanced molecular ionic and net ionic equations nie for each of the following reactions. Net ionic equation worksheet.

### Net Ionic Equation Worksheet Answers - Thekidsworksheet

Molecular, complete and net ionic equations worksheet Write the balanced molecular, complete ionic, and net ionic equations for each of the following reactions. Assume all reactions occur in aqueous solution. (Use the solubility rules when needed.)

### Net Ionic Equation Worksheet - Sophie B. Wright Charter School

WRITING NET IONIC EQUATIONS . 1. First write out the proper formulas for the reactants. Ex: Aqueous solutions of sodium phosphate and ruthenium(VI) nitrate are reacted to form a precipitate. Sodium phosphate sodium is always  $\text{Na}^+$ , phosphate is one of the polyatomics (that you're supposed to have memorized by now) and has the formula of  $\text{PO}_4^{3-}$

### Writing net ionic equations - New Providence School ...

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View Net Ionic Equations Worksheet.docx from CHM 112 at Pulaski County High School. Net Ionic Equations Worksheet Write balanced molecular, ionic, and net ionic equations (NIE) for each of the

### **Net Ionic Equations Worksheet.docx - Net Ionic Equations ...**

Net Ionic Equation Worksheet . READ THIS: When two solutions of ionic compounds are mixed, a solid may form. This type of reaction is called a precipitation reaction, and the solid produced in the reaction is known as the precipitate.

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