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Organometallic chemistry is the study of organometallic compounds, chemical compounds containing at least one chemical bond between a carbon atom of an organic molecule and a metal, including alkaline, alkaline earth, and transition metals, and sometimes broadened to include metalloids like boron, silicon, and tin, as well. Aside from bonds to organyl fragments or molecules, bonds to 'inorganic' carbon, like carbon monoxide, cyanide, or carbide, are generally considered to be organometallic as w

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Solvents featuring ether functionalities constitute highly valuable media for performing chemistry with organometallic reagents []. This is mainly due to their capability to disaggregate such species, thus modulating their reactivity []. Unfortunately, common solvents such as tetrahydrofuran (THF) tend to react with highly basic carbanions, thus requiring the employment of low temperatures to ...

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I. Basic Principles A. Organometallic Mechanisms Oxidation State: The oxidation state of a metal is defined as the charge left on the metal after all ligands have been removed in their natural, closed-shell configuration. This is a formalism and not a physical property! d-Electron Configuration: position in the periodic table minus oxidation state.

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Organometallic chemistry refers to that branch of chemistry that studies the chemical compounds, which have one bond between carbon atom and a metal. It includes the study of alkaline Earth, alkaline and transition metal. This subject uses the elements of both organic and inorganic chemistry.

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