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1) each element is made of atoms 2) all atoms of a given element are identical there are isotopes 3) atoms are neither created nor destroyed 4) atoms are more than 1 element form a compound

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Name: _____Ella_____ Study Guide Test 1: Subatomic Particles
CHM.1.1.1 Analyze the structure of atoms, isotopes, and ions.
CHM.1.1.2 Analyzed an atom in terms of the location of electrons. 1. Describe how to write a chemical symbol. The first letter is capitalized and if needed the second letter is lower cases. 2. Complete the chart.

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Basic Building Blocks of Matter. Atom, element, compound.
Atom. An atom is the smallest unit of an element that maintains the chemical identity of that element. Element. An element is a

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pure substance that cannot be broken down into simpler, stable substances and is made of one type of atom. Compound.

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Molecules: the smallest unit of a substance that keeps all of the physical and chemical properties of that substance; it can consist of one atom or two or more atoms bounded together. (usually consists of two or more atoms combined in a definite ratio).

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Unit 1 Chemistry Study Guide What characterizes a prokaryotic cell? Why do experiments need a control? Why do scientists use the scientific method to study environmental problems? List the hierarchy of organization for living things & for classification.

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Step 1: Write the unbalanced equation: $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
Step 2: Apply Avogadro's law: $40 \text{ cm}^3 \text{ ' } 20 \text{ cm}^3 \text{ nd } 2$
volumes 1 volume 2 mole(cule)s 1 mole(cule)
Step 3: Write the equation: $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l})$

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1. Multiply each isotope's mass by its abundance. If your abundance is a percent, divide your answer by 100. 2. Add these values together.

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(1 pound = 453.6 g) Convert 1.83 m to inches. (1 cm = 2.54 in)
8. Be able to calculate density or use density to find mass or volume. $D = m/vol$; typical units = g/mL or g/cm³ Calculate the density in g/mL or g/cm³ given mass = 234 g and volume = 2.2 cm³. Calculate the density in g/mL or g/cm³ given mass = 1.2 lbs and volume = 2.2 ft³. (1 lb =

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Chemistry Unit 1 Study Guide □□question Describe three states of matter answer Three states of matter: solid, liquid and gas.

-Solid: constant shape, constant volume, particles bound

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pool guy, chef, make up person, and chemistry teacher 2. The

scientific method is a process of studying natural phenomenons

Steps: First

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...

Chemistry A Unit 1 Study Guide Students, The Unit 1 Test is a

timed assessment. You will have 55 minutes to complete the

test. There are 13 multiple-choice questions and 6 short answer

questions. Below are the vocabulary and major concepts you

need to understand before taking the exam.

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The atom, along with its protons, neutrons, and electrons, is the basic unit of matter and serves as our starting point for the study of chemistry. Scientists have studied atoms for hundreds of years and have developed a number of different models to describe them.

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